

Why aren't the elements listed in alphabetical order?

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As charts go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. Those properties relate to the electronic structure of the atoms of each element. In today's activity, you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

A blank periodic table grid with the following labels:

- s-block**: Located above the first two columns (Groups 1 and 2).
- p-block**: Located above the last six columns (Groups 13 to 18).
- d-block**: Located above the central ten columns (Groups 3 to 10).
- f-block**: Located below the main grid, centered under the d-block.

The grid has 7 rows and 18 columns. The first two columns are labeled 's-block', the next ten columns are labeled 'd-block', and the last six columns are labeled 'p-block'. The f-block is shown as a separate section below the main grid, consisting of two rows of 14 columns each.

1. Obtain a card with electron configurations for your group as assigned by your teacher.
 - a. Record your team number_____ and team name _____
 - b. For each element on your card, write the end of the configuration (the underlined portion).
2. What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

Cracking The Periodic Table Code

RS Peters



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