### Determining the Stoichiometry of Chemical Reactions continued

## Pre-Laboratory Assignment

The following values were obtained in a continuous variations experiment designed to find the coefficients in the equation for the reaction between 0.5 M solutions of AgNO<sub>3</sub> and K<sub>2</sub>CrO<sub>4</sub>. One of the products is a precipitate:

Experiment	mL AgNO,	THOO	
		mL K <sub>2</sub> CrO <sub>4</sub>	Grams Precipitate
1	5.0	45.0	1.7
2	15.0	35.0	5.0
3	25.0	25.0	8.3
4	30.0	20.0	10.0
5	35.0	15.0	9.9
6	40.0	10.0	6.6
7	45.0	5.0	3.3

Plot the mL of AgNO<sub>3</sub> versus grams precipitate on graph paper. Label axes and space the data so that the graph reflects the precision of the values given. Use a ruler to draw two best-fitting straight lines through the increasing and decreasing data points. Determine the stoichiometry of the reaction from the intersection of these lines.

# **Determining The Stoichiometry Of Chemical Reactions Answers**

**G** Orfield

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web 2 gilgal the place of beginning 2023 07 22 remain determined and focused for in due season you shall reap if you faint not gal 6 9 the song of the sea peter lang we are not likely to discover the treasure of the so called old testament if we regard whole passages as historically interesting yet without relevance to our