

General Chemistry Reference Sheet

This reference sheet addresses some of the more peculiar pieces of information that need to be memorized in a general chemistry course. It also contains a simple set of essential formulas in chemistry with cautions, explanations, and general tips.

This sheet is meant to be as concise as possible, and many informations in the textbook is left out in favor of cautions and tips. This sheet is, therefore, best used as a supplement to, not a replacement of, the textbook.

SI Fundamental Units

Mass	Kilogram (kg)
Length	Meter (m)
Time	Second (s)
Temperature	Kelvin (K)
Amount of substance	Mole (mol)
Electric current	Ampere (A)
Luminous intensity	Candela (cd)

Atomic Experiments and Models

J. J. Thomson	Discovered $e^{-}$ ; Cathode ray
	Plum pudding model
R. A. Millikan	Measured charge of $e^{-}$ ; Oil drop
H. Becquerel/M. Curie	Discovered radioactivity
E. Rutherford	Discovered $\alpha$ , $\beta$ , and $\gamma$ rays
	Discovered nucleus; Gold foil experiment
J. Chadwick	Discovered neutrons
N. Bohr	Bohr model (electron orbits)
Quantum mechanists	Quantum model

Polyatomic Ions

$\text{NH}_4^{+}$	ammonium	$\text{OH}^{-}$	hydroxide
$\text{CN}^{-}$	cyanide	$\text{C}_2\text{O}_4^{2-}$	oxalate
$\text{O}_2^{2-}$	peroxide	$\text{CNO}^{-}$	cyanate
$\text{HSO}_3^{-}$	hydrogen sul- fate	$\text{C}_2\text{H}_3\text{O}_2^{-}$	acetate
$\text{SCN}^{-}$	thiocyanate	$\text{NO}_3^{-}$	nitrate
$\text{SO}_3^{2-}$	sulfite	$\text{ClO}_4^{-}$	perchlorate
$\text{CO}_3^{2-}$	carbonate	$\text{ClO}_3^{-}$	chlorate
$\text{PO}_4^{3-}$	phosphate	$\text{ClO}^{-}$	hypochlorite
$\text{S}_2\text{O}_4^{2-}$	dithionite	$\text{HPO}_4^{2-}$	hydrogen phosphate
$\text{CrO}_4^{2-}$	chromate	$\text{H}_2\text{O}^{+}$	hydronium
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	$\text{PO}_3^{2-}$	phosphite
$\text{MnO}_4^{-}$	permanganate	$\text{Hg}_2^{2+}$	mercury (I)
$\text{N}_3^{-}$	azide	$\text{C}_4^{4-}$	carbide
$\text{C}_4\text{H}_4\text{O}_6^{2-}$	tartarate	$\text{S}_2^{2-}$	dissulfide
$\text{O}_2^{-}$	superoxide	$\text{AsO}_3^{3-}$	arsenite
$\text{PO}_3^{3-}$	hypophosphite	$\text{AsO}_4^{3-}$	arsenate
$\text{SO}_3^{2-}$	sulfite	$\text{P}_3\text{O}_6^{4-}$	pyrophosphate

Ionic Solubility Chart

Soluble	Exceptions
$\text{NO}_3^{-}$	None
$\text{CH}_3\text{COO}^{-}$	None
$\text{Cl}^{-}$	$\text{Ag}^{+}$ , $\text{Hg}_2^{2+}$ , $\text{Pb}^{2+}$
$\text{Br}^{-}$	$\text{Ag}^{+}$ , $\text{Hg}_2^{2+}$ , $\text{Pb}^{2+}$
$\text{I}^{-}$	$\text{Ag}^{+}$ , $\text{Hg}_2^{2+}$ , $\text{Pb}^{2+}$
$\text{SO}_4^{2-}$	$\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Hg}_2^{2+}$ , $\text{Pb}^{2+}$
Insoluble	Exceptions
$\text{S}^{2-}$	$\text{NH}_4^{+}$ , alkali metal cations, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$
$\text{CO}_3^{2-}$	$\text{NH}_4^{+}$ , alkali metal cations
$\text{PO}_4^{3-}$	$\text{NH}_4^{+}$ , alkali metal cations
$\text{OH}^{-}$	$\text{NH}_4^{+}$ , alkali metal cations, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$

Strong Acids and Bases

Strong acids and bases dissociate in water completely.

Strong Acids		Strong Bases
HCl	HClO <sub>4</sub>	Alkali metal hydroxides
HBr	HNO <sub>3</sub>	Ca(OH) <sub>2</sub>
HI	H <sub>2</sub> SO <sub>4</sub>	Sr(OH) <sub>2</sub>
HClO <sub>2</sub>	Ba(OH) <sub>2</sub>	

Activity Series

Metals below  $\text{H}^{+}$  cannot react with acids to form  $\text{H}_2$ . More active metals are better reducing agents.

From most active to least active:

$\text{Li}^{+}$ ,  $\text{K}^{+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Na}^{+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Cr}^{3+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Sn}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{H}^{+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Ag}^{+}$ ,  $\text{Hg}_2^{2+}$ ,  $\text{Pt}^{2+}$ ,  $\text{Au}^{3+}$

Flame Colors

Calcium	Brick red
Copper (I)	Blue
Copper (II)	Green or blue-green
Potassium	Libre
Lithium	Dark red
Sodium	Bright yellow
Strontium	Red
Barium	Light green
Iron (III)	Gold
Cesium	Blue-Violet
Iodine	Blue
Lead	Blue
Rubidium	Red-Violet

Phase Changes

	From solid	From liquid	From gas
To solid	-	freezing	deposition
To liquid	melting	-	condensation
To gas	sublimation	evaporation	-

Solution Colors

Copper (II)	Blue
Nickel	Green
Permanganate	Purple
Chromate	Yellow
Dichromate	Orange
Iron (II)	Light blue
Iron (III)	Rusty yellow

Thermodynamic Laws

**First Law:** Energy cannot be created nor destroyed. It can only be transferred in the form of either heat or work.

**Second Law:** Any spontaneous reaction increases the entropy of the universe.

**Third Law:** An ideal solid crystal at 0 K has an entropy of 0.

Thermodynamic Formulas

Standard thermodynamic conditions:	298 K; 1 atm; 1 M
Kinetic energy	$K = mv^2/2$
Electrostatic potential energy	$U_E = (k_e Q_1 Q_2)/d$
Internal energy	$\Delta E = q + w$
Enthalpy	$H = E + PV$
Specific heat	$s = q/(m \cdot \Delta T)$
Entropy in reversible reaction	$\Delta S_{\text{system}} = (\Delta H)/T$ $\Delta S_{\text{surrounding}} = -(\Delta H)/T$
Microstate-entropy relationship	$S = k \ln W$
Gibbs free energy	$G = H - TS$
Gibbs free energy change	$\Delta G = \Delta H - T \Delta S$ $\Delta G = \Delta G^{\circ} + RT \ln Q$ $\Delta H_{\text{total}} = T \Delta S_{\text{total}}$
Hess's Law	

Constants

Boltzmann's constant	$k_B = 1.381 \times 10^{-23} \text{ m}^2\text{kg} \cdot \text{s}^{-2}\text{K}^{-1}$
Coulomb's constant $k_C = 1/(4\pi\epsilon_0)$	$= 8.988 \times 10^9 \text{ J} \cdot \text{m}^2/\text{C}^2$
Avogadro's number	$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	$F = 9.649 \times 10^4 \text{ C/mol}$
Planck's constant	$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$
Ideal gas constants	$R = 0.0821 \text{ (L} \cdot \text{atm)/(mol} \cdot \text{K)}$ $R = 8.314 \text{ J/(mol} \cdot \text{K)}$
Vacuum permittivity $\epsilon_0 = 1/(\mu_0 c^2)$	$= 8.854 \times 10^{-12} \text{ F/m}$
Vacuum permeability	$\mu_0 = 1.257 \times 10^{-6} \text{ N}\cdot\text{A}^{-2}$
Atomic mass	1 amu = 1.661 $\times 10^{-24}$ g
Electron charge	$e = 1.602 \times 10^{-19}$ C
Electronvolt	1 eV = 1.602 $\times 10^{-19}$ J
Atmospheric pressure	1 atm = 1.013 $\times 10^5$ Pa
Absolute zero	0 K = -273.15 $^{\circ}\text{C}$
Speed of light in vacuum	$c = 2.998 \times 10^8 \text{ m/s}$

Quantum Mechanical Formulas

Energy of a quantum	$E = hf$
Wavelength-frequency relationship	$c = \nu \cdot \lambda$
Probability distribution	$P_V = \iiint  \psi(x,y,z) ^2 dx dy dz$

# Cst Chemistry Reference Sheet

**Alan F. Cassel, Berinda P. Johnson**



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*PDQ Primary Reference Guide*, 1988      Alternatives for High-Level Waste Salt Processing at the Savannah River Site  
National Research Council, Board on Chemical Sciences and Technology, Board on Radioactive Waste  
Management, Committee on Cesium Processing Alternatives for High-Level Waste at the Savannah River Site, 2000-10-30 The  
Second World War introduced the world to nuclear weapons and their consequences Behind the scene of these nuclear  
weapons and an aspect of their consequences is radioactive waste Radioactive waste has varying degrees of harmfulness and  
poses a problem when it comes to storage and disposal Radioactive waste is usually kept below ground in varying containers  
which depend on how radioactive the waste is High level radioactive waste HLW can be stored in underground carbon steel  
tanks However radioactive waste must also be further immobilized to ensure our safety There are several sites in the United  
States where high level radioactive waste HLW are stored including the Savannah River Site SRS established in 1950 to  
produce plutonium and tritium isotopes for defense purposes In order to further immobilize the radioactive waste at this site  
an in tank precipitation ITP process is utilized Through this method the sludge portion of the tank wastes is being removed  
and immobilized in borosilicate glass for eventual disposal in a geological repository As a result a highly alkaline salt present  
in both liquid and solid forms is produced The salt contains cesium strontium actinides such as plutonium and neptunium and  
other radionuclides But is this the best method The National Research Council NRC has empanelled a committee at the  
request of the U S Department of Energy DOE to provide an independent technical review of alternatives to the discontinued  
in tank precipitation ITP process for treating the HLW stored in tanks at the SRS Alternatives for High Level Waste Salt  
Processing at the Savannah River Site summarizes the finding of the committee which sought to answer 4 questions  
including Was an appropriately comprehensive set of cesium partitioning alternatives identified and are there other  
alternatives that should be explored and Are there significant barriers to the implementation of any of the preferred  
alternatives taking into account their state of development and their ability to be integrated into the existing SRS HLW  
system      *Book of ASTM Standards, with Related Material* American Society for Testing and Materials, 1976      **Reader's**  
**Index and Guide**, 1917      Research Reporting Series, 1979      **EPA-600/2**, 1979      *Annual Book of ASTM Standards*  
American Society for Testing and Materials, 1978      Tables of Molecular Vibrational Frequencies Takehiko  
Shimanouchi, 1972      Evaluation of Dewatering Devices for Producing High-solids Sludge Cake Alan F. Cassel, Berinda P.  
Johnson, 1979      **Next Generation Kinase Inhibitors** Paul Shapiro, 2020-07-14 Protein kinases are fascinating enzymes  
that maintain the proper function of nearly every task performed by the cells of the human body By extracting a phosphate  
from the energy molecule ATP and linking it to another protein protein kinases alter the structure and ultimate function of  
other proteins In this way protein kinases help monitor the extracellular environment and integrate signaling cues that for  
the most part are beneficial for human health and survival However protein kinases are often dysregulated and responsible

for the initiation and progression of many types of cancers inflammatory disorders and other diseases Thus decades of research have revealed much about how protein kinases are regulated and approaches to inhibit these enzymes to treat disease However nearly 30 years since the identification of the first clinically beneficial small molecule protein kinase inhibitor there are only a few examples where these drugs provide sustained and durable patient responses The goal of this book is to provide biomedical scientists graduate and professional degree students insight into different approaches using small molecules to block specific protein kinase functions that promote disease     The Reader's Index Croydon Public Libraries,1917     **Sustainable Solvents** James H Clark,Andrew Hunt,Corrado Topi,Giulia Paggiola,James

Sherwood,2017-05-11 Exploring the solvent life cycle this book provides a critical assessment on emerging solvents for academics and industrialists working with solvents as well as policy makers     **NBS Special Publication** ,1968

**Reference Catalogue of Current Literature** ,1902     The Chemical News ,1917     *Synthetics, Mineral Oils, and Bio-Based Lubricants* Leslie R. Rudnick,2013-02-04 Highlighting the major economic and industrial changes in the lubrication industry since the first edition Synthetics Mineral Oils and Bio Based Lubricants Second Edition outlines the state of the art in each major lubricant application area Chapters cover trends in the major industries such as the use of lubricant fluids growth or decl     **Clinical Diagnosis by Laboratory Methods** James Campbell Todd,1974 1999 Gift of Forde A McIver M D     **The New Britannica/Webster Dictionary & Reference Guide** Encyclopaedia Britannica, Inc,1981

**New Serial Titles** ,1992 A union list of serials commencing publication after Dec 31 1949     *The Library Catalogs of the Hoover Institution on War, Revolution, and Peace, Stanford University* Hoover Institution on War, Revolution, and Peace,1969

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