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## Chemical Engineering Thermodynamics (SKKK 2213)

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Chapter 2

### 2.1 (7<sup>th</sup> edition Prob. 2.1)

(a) The work done on the water is equal to the work done by gravity on the weight, which is equal to force times distance. The force is  $mg = 35 \text{ kg} \cdot 9.8 \text{ m s}^{-2} = 343 \text{ N}$ . The distance is 5 m, so the work is  $343 \text{ N} \cdot 5 \text{ m} = 1715 \text{ N m} = 1715 \text{ J}$ .

(b) The internal energy change is equal to the work done, 1715 J.

(c) The internal energy change (1715 J) is equal to the total heat capacity times the temperature change ( $\Delta U = mC_p \Delta T$ ) or  $1715 \text{ J} = 25 \text{ kg} \cdot 4180 \text{ J kg}^{-1} \text{ K}^{-1} \Delta T$ , from which

$$\Delta T = 1715 / (25 \cdot 4180) = 0.016 \text{ K}$$

So, the final temperature of the water is  $20.02^\circ\text{C}$  (the initial temperature was  $20^\circ\text{C}$ ).

(d) To return the water to its original state, one must remove 1715 J as heat. (e) Zero, zero, zero! The total energy change of the universe is always zero (in the absence of nuclear reactions, of course).

### 2.2 (7<sup>th</sup> edition Prob. 2.2)

The work done on the water is equal to the work done by gravity on the weight, which is equal to force times distance. The force is  $mg = 35 \text{ kg} \cdot 9.8 \text{ m s}^{-2} = 343 \text{ N}$ . The distance is 5 m, so the work is  $343 \text{ N} \cdot 5 \text{ m} = 1715 \text{ N m} = 1715 \text{ J}$ .

a) The internal energy change of the system (water plus container) is equal to the work done, 1715 J. However, not all of this is change in the internal energy of the water. The heat capacity of the water is  $(25/30) = 5/6$  of the total heat capacity. Thus, when the temperature of the water and container both increase,  $5/6$  of the energy goes into the water, and  $1/6$  goes into the container. So, the internal energy of the water increases by  $5/6 \cdot 1715 = 1429 \text{ J}$ . The internal energy of the container changes by 286 J.

b) The internal energy change (1715 J) is equal to the total heat capacity times the temperature change ( $\Delta U = mC_p \Delta T$ ) or  $1715 \text{ J} = 30 \text{ kg} \cdot 4180 \text{ J kg}^{-1} \text{ K}^{-1} \Delta T$ , from which

$$\Delta T = 1715 / (30 \cdot 4180) = 0.0137 \text{ K}$$

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**H. Scott Fogler**



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