

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

### Density Practice Problems **KEY**

This image will help you in figuring out how to solve density problems:

$$\frac{M}{dV}$$

Simply cover up whichever value you need to calculate and the other two are shown in their proper placement, be it to multiply or to divide.

For example, cover up the M. This leave you with dV (ignore the fact that it is in the denominator). Density times volume will give you mass. You can also check it out by way of the units: (g / cm<sup>3</sup>) x cm<sup>3</sup> cancels out the volume unit leaving grams, the desired unit for mass.

*Solve the following problems **showing all your work** including equations and units.*

1) A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its **density**?

$$D = m / v$$

$$D = 40.5 \text{ g} / 15 \text{ ml} = \mathbf{2.7 \text{ g/mL}}$$

2) Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury used to fill the cylinder weighs 306.0 g. From this information, calculate the **density** of mercury.

$$D = m / v$$

$$D = 306.0 \text{ g} / 22.5 \text{ mL} = \mathbf{13.6 \text{ g/mL}}$$

3) What is the **mass** of the ethyl alcohol that exactly fills a 200.0 mL container? The density of ethyl alcohol is 0.789 g/mL.

$$m = D \times V$$

$$m = 0.789 \text{ g/mL} \times 200.0 \text{ mL} = \mathbf{157.8 \text{ g}}$$

4) A rectangular block of copper metal weighs 1896 g. The dimensions of the block are 8.4 cm by 5.5 cm by 4.6 cm. From this data, what is the **density** of copper?

$$V = L \times W \times H$$

$$V = (8.4 \text{ cm}) (5.5 \text{ cm}) (4.6 \text{ cm}) = 212.52 \text{ cm}^3$$

$$D = m / V$$

$$D = 1896 \text{ g} / 212.52 \text{ cm}^3 = \mathbf{8.9 \text{ g/cm}^3}$$

# Density Practice Problems With Answers Fordhamprep

**IM Harris**



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