

Density and Specific Gravity

3

A. Density of Liquids

A.1 Density of Water

1. Volume (mL)
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 20.0 \text{ mL} \\
 0.00 \text{ g} \\
 15.82 \text{ g} \\
 \hline
 20.420 \text{ g} \\
 D = \frac{m}{V} = \frac{20.420}{20.0} = 1.01 \text{ g/mL}
 \end{array}$$

A.2 Density of an Unknown Liquid

1. Type of liquid sample
Volume of liquid
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 \text{Salt water} \\
 20.1 \text{ mL} \\
 0.00 \text{ g} \\
 15.89 \text{ g} \\
 \hline
 21.24 \text{ g} \\
 \frac{21.24 \text{ g}}{20.1 \text{ mL}} = 1.05 \text{ g/mL}
 \end{array}$$

B. Specific Gravity

1. Specific gravity of water
2. Specific gravity of liquid sample
3. Specific gravity of water using a hydrometer
4. Specific gravity of liquid using a hydrometer

$$\begin{array}{r}
 1.01 \\
 1.05 \\
 0.995 \\
 0.790
 \end{array}$$

How do the *calculated* specific gravity of water and the liquid sample (1, 2) compare to the hydrometer readings (3, 4) for water and the liquid sample?

Density And Specific Gravity Laboratory Manual

Edward Leamington Nichols



Density And Specific Gravity Laboratory Manual:

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Table of Contents Density And Specific Gravity Laboratory Manual

1. Understanding the eBook Density And Specific Gravity Laboratory Manual
 - The Rise of Digital Reading Density And Specific Gravity Laboratory Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Density And Specific Gravity Laboratory Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Density And Specific Gravity Laboratory Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Density And Specific Gravity Laboratory Manual
 - Personalized Recommendations
 - Density And Specific Gravity Laboratory Manual User Reviews and Ratings
 - Density And Specific Gravity Laboratory Manual and Bestseller Lists
5. Accessing Density And Specific Gravity Laboratory Manual Free and Paid eBooks
 - Density And Specific Gravity Laboratory Manual Public Domain eBooks
 - Density And Specific Gravity Laboratory Manual eBook Subscription Services
 - Density And Specific Gravity Laboratory Manual Budget-Friendly Options

6. Navigating Density And Specific Gravity Laboratory Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Density And Specific Gravity Laboratory Manual Compatibility with Devices
 - Density And Specific Gravity Laboratory Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Density And Specific Gravity Laboratory Manual
 - Highlighting and Note-Taking Density And Specific Gravity Laboratory Manual
 - Interactive Elements Density And Specific Gravity Laboratory Manual
8. Staying Engaged with Density And Specific Gravity Laboratory Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Density And Specific Gravity Laboratory Manual
9. Balancing eBooks and Physical Books Density And Specific Gravity Laboratory Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Density And Specific Gravity Laboratory Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Density And Specific Gravity Laboratory Manual
 - Setting Reading Goals Density And Specific Gravity Laboratory Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Density And Specific Gravity Laboratory Manual
 - Fact-Checking eBook Content of Density And Specific Gravity Laboratory Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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