

WRITTEN WORK NO. 1

Thermochemistry

Instruction: Solve the following items below.

1. Determine the heat of formation of liquid hydrogen peroxide, H_2O_2 at 25°C from the following thermochemical equations:



$$\Delta H = -241.82 \text{ kJ/mol}$$



$$\Delta H = -926.92 \text{ kJ/mol}$$



$$\Delta H = -1,070.60 \text{ kJ/mol}$$



$$\Delta H = -498.34 \text{ kJ/mol}$$



$$\Delta H = +52.46 \text{ kJ/mol}$$

Solution:

Final Answer:

2. Predict the sign of ΔS for the following processes and explain why.



3. For a process, $\Delta H_{\text{system}} = 145 \text{ kJ}$ and $\Delta S_{\text{system}} = 322 \text{ J/K}$, is the process spontaneous at 382K ?

Solution:

Final Answer:



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IM Harris



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