

Activity C:
Calculating
specific heat

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Introduction: The specific heat capacity of a substance is the amount of energy needed to change the temperature of that substance by 1 °C. Specific heat capacity can be calculated using the following equation:
$$q = mc\Delta T$$

In the equation q represents the amount of heat energy gained or lost (in **joules**), m is the mass of the substance (in grams), c is the specific heat capacity of the substance (in J/g °C), and ΔT is the temperature change of the substance (in °C).

Goal: Calculate the specific heat capacities of copper, granite, lead, and ice.

1. **Solve:** When you mix two substances, the heat gained by one substance is equal to the heat lost by the other substance. Suppose you place 125 g of aluminum in a calorimeter with 1,000 g of water. The water changes temperature by 2 °C and the aluminum changes temperature by -74.95 °C.

- A. Water has a known specific heat capacity of 4.184 J/g °C. Use the specific heat equation to find out how much heat energy the water gained (q).

$$\Delta T Q = 1000 \times 4.18 \times 2Q = 8360 \text{ J} \quad 8360 \text{ J}; 8.4 \text{ kJ}$$

- B. Assume that the heat energy gained by the water is equal to the heat energy lost by the aluminum. Use the specific heat equation to solve for the specific heat of aluminum. (Hint: Because heat energy is lost, the value of q is negative.)

0.893

Aluminum's accepted specific heat value is 0.900 J/g °C. Use this value to check your work.

2. **Calculate:** Use the Gizmo to mix 200 g of copper at 100 °C with 1,000 g of water at 20 °C.

- A. What is the final temperature? 21.42
- B. Calculate the temperature change of each substance by subtracting the initial temperature from the final temperature.

$\Delta T_{\text{water}} =$ 1.49

$\Delta T_{\text{copper}} =$ 20

- C. How much heat energy (q) did the water gain? 18
- D. Now solve for the specific heat (c) of copper; it is 20°C

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A Gutmann



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Table of Contents Explore Learning Calorimetry Lab Answer Key

1. Understanding the eBook Explore Learning Calorimetry Lab Answer Key
 - The Rise of Digital Reading Explore Learning Calorimetry Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Calorimetry Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Calorimetry Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Calorimetry Lab Answer Key
 - Personalized Recommendations
 - Explore Learning Calorimetry Lab Answer Key User Reviews and Ratings
 - Explore Learning Calorimetry Lab Answer Key and Bestseller Lists
5. Accessing Explore Learning Calorimetry Lab Answer Key Free and Paid eBooks
 - Explore Learning Calorimetry Lab Answer Key Public Domain eBooks
 - Explore Learning Calorimetry Lab Answer Key eBook Subscription Services
 - Explore Learning Calorimetry Lab Answer Key Budget-Friendly Options

6. Navigating Explore Learning Calorimetry Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Calorimetry Lab Answer Key Compatibility with Devices
 - Explore Learning Calorimetry Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Calorimetry Lab Answer Key
 - Highlighting and Note-Taking Explore Learning Calorimetry Lab Answer Key
 - Interactive Elements Explore Learning Calorimetry Lab Answer Key
8. Staying Engaged with Explore Learning Calorimetry Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Calorimetry Lab Answer Key
9. Balancing eBooks and Physical Books Explore Learning Calorimetry Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Calorimetry Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Calorimetry Lab Answer Key
 - Setting Reading Goals Explore Learning Calorimetry Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Calorimetry Lab Answer Key
 - Fact-Checking eBook Content of Explore Learning Calorimetry Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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by guest perfection dzogchen both of these experimented with sensory deprivation and developed yogas involving long periods of dwelling in dark rooms or gazing at the open sky these produced unusual experiences