

## WRITTEN WORK NO. 1

**Thermochemistry**

Instruction: Solve the following items below.

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1. Determine the heat of formation of liquid hydrogen peroxide,  $\text{H}_2\text{O}_2$  at  $25^\circ\text{C}$  from the following thermochemical equations:



$$\Delta H = -241.82 \text{ kJ/mol}$$



$$\Delta H = -926.92 \text{ kJ/mol}$$



$$\Delta H = -1,070.60 \text{ kJ/mol}$$



$$\Delta H = -498.34 \text{ kJ/mol}$$



$$\Delta H = +52.46 \text{ kJ/mol}$$

Solution:

Final Answer:

2. Predict the sign of  $\Delta S$  for the following processes and explain why.



3. For a process,  $\Delta H_{\text{system}} = 145 \text{ kJ}$  and  $\Delta S_{\text{system}} = 322 \text{ J/K}$ , is the process spontaneous at  $382\text{K}$ ?

Solution:

Final Answer:



## Gc2 Worksheet 6

**Siddappa N.Byrareddy**



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## **Table of Contents Gc2 Worksheet 6**

1. Understanding the eBook Gc2 Worksheet 6
  - The Rise of Digital Reading Gc2 Worksheet 6
  - Advantages of eBooks Over Traditional Books
2. Identifying Gc2 Worksheet 6
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gc2 Worksheet 6
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gc2 Worksheet 6
  - Personalized Recommendations
  - Gc2 Worksheet 6 User Reviews and Ratings
  - Gc2 Worksheet 6 and Bestseller Lists
5. Accessing Gc2 Worksheet 6 Free and Paid eBooks
  - Gc2 Worksheet 6 Public Domain eBooks
  - Gc2 Worksheet 6 eBook Subscription Services
  - Gc2 Worksheet 6 Budget-Friendly Options
6. Navigating Gc2 Worksheet 6 eBook Formats
  - ePub, PDF, MOBI, and More
  - Gc2 Worksheet 6 Compatibility with Devices
  - Gc2 Worksheet 6 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gc2 Worksheet 6
  - Highlighting and Note-Taking Gc2 Worksheet 6
  - Interactive Elements Gc2 Worksheet 6
8. Staying Engaged with Gc2 Worksheet 6

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Gc2 Worksheet 6
- 9. Balancing eBooks and Physical Books Gc2 Worksheet 6
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gc2 Worksheet 6
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gc2 Worksheet 6
  - Setting Reading Goals Gc2 Worksheet 6
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gc2 Worksheet 6
  - Fact-Checking eBook Content of Gc2 Worksheet 6
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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