

1. For **each** of the following sets of sublevels, determine which has the lowest energy, and which has the highest energy.

(a) 2s, 3s, 4s

(b) 4f, 4p, 4d

(c) 4d, 4f, 3d

(d) 6s, 3d, 5f

Ex for (a) - Lowest 2s, highest 4s

2. Explain what  $6s^2$  means.

3. What is the maximum number of electrons that can be found in an s – sublevel? \_\_\_\_ p – sublevel? \_\_\_\_ d – sublevel? \_\_\_\_ f – sublevel? \_\_\_\_

4. Indicate which of these (there may be more than one) orbital designations is incorrect:

(a) 4s

(b) 3f

(c) 2d

(d) 3d

5. Draw energy level diagrams for

(a) sodium

(b) silicon

(c) beryllium

(d) cobalt

6. How many unpaired electrons are there in each of the atoms in question 5?

7. Give the electron configuration for:

(a) lithium

(b) oxygen

(c) selenium

(d) iodine

(e) silver

(f) arsenic

(g) indium

(h) potassium

8. Give the name and symbol for the atom whose electron configuration corresponds to each of the following electron configurations.

a)  $1s^2 2s^2 2p^6 3s^2 3p^5$

b)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$

c)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^6$

d)  $1s^2 2s^2 2p^6 3s^2 3p^2$

e)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$

9. Write the shorthand notation for the above electron configurations.

# Electron Configuration Exam Questions

**Judith Potter, Peter Cann**



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