

Concentration and Molarity II: Dilution and Evaporation PhET-Chemistry Lab



Introduction: In the first part of this lab, you learned about the actions of solid salts added to water to form solutions of various concentrations. In this section, you will use a powerful and simple formula to determine the concentration of a solution when it has been diluted with additional water or made more concentrated due to evaporation of water.

Some handy equations for you to define:

Molarity _____

Dilution _____

Evaporation _____

Procedure: PHET-Poly with the Chem II-Chemistry II Concentration 100%

$$M_1V_1 = M_2V_2$$

Part I: Dilution

At this point you should be familiar with the Concentration simulation. Because some time may have passed, take a few minutes to re-learn the simulation. Pay particular attention to the effect of evaporation and addition of water to the solution's concentration.



How does the concentration change as additional water is added? _____

Why? _____



How does the concentration change as evaporation occurs? _____

Why is this? _____

Does evaporation change the concentration of a saturated solution? _____

Why is this? _____



Using the concentrated solution spigot, add a 10 Liter of Drink Mix to an empty beaker. What is the concentration?

_____ Is this solution saturated? _____ How do you know? _____

Fill the beaker with another 10 Liter of water. What is the new concentration? _____

Complete the table below, using 100% Cobalt chloride in an empty beaker, writing the concentration in the boxes provided.

Dilute .25L of spigot solution	.25L spigot+.25L water	.25L spigot + .50 L water	.25L spigot + .75 L water

Repeat the exercises, using 100% Potassium dichromate in an empty beaker.

Dilute .25L of spigot solution	.25L spigot+.25L water	.25L spigot + .50 L water	.25L spigot + .75 L water

What do you notice about the concentration change as each addition of .25L of water is added to the concentrated spigot solution? _____

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Answers Key

Victor M. Corman



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