

**Assessment #3 Review**

Name: \_\_\_\_\_

You must show all work for credit.

Due Date: \_\_\_\_\_ (No late packets will be accepted)

**I. Balance Equations***Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of  $\text{Fe}_2\text{O}_3$  in reaction #6? 3

*Write the balanced equation for problems # 8 and 9:*

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

moles

12. Using the balanced equation shown in problem #10, list the products of the reaction.

$\text{NH}_3$

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

$\text{N}_2$ ,  $\text{H}_2$

**II. Types of Reactions***Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



# Final Chemistry Exam Review Sheet With Answers

**Jiben Roy**



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Roald Hoffmann on the Philosophy, Art, and Science of Chemistry Jeffrey Kovac, Michael

Weisberg, 2012-02-01 Nobel laureate Roald Hoffmann's contributions to chemistry are well known Less well known however is that over a career that spans nearly fifty years Hoffmann has thought and written extensively about a wide variety of other topics such as chemistry's relationship to philosophy literature and the arts including the nature of chemical reasoning the role of symbolism and writing in science and the relationship between art and craft and science In Roald Hoffmann on the Philosophy Art and Science of Chemistry Jeffrey Kovac and Michael Weisberg bring together twenty eight of Hoffmann's most important essays Gathered here are Hoffmann's most philosophically significant and interesting essays and lectures many of which are not widely accessible In essays such as Why Buy That Theory Nearly Circular Reasoning How Should Chemists Think The Metaphor Unchained Art in Science and Molecular Beauty we find the mature reflections of one of America's leading scientists Organized under the general headings of Chemical Reasoning and Explanation Writing and Communicating Art and Science Education and Ethics these stimulating essays provide invaluable insight into the teaching and practice of science

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Sloan, Clara Hurd, 2014-07-31 Includes a CD ROM with 6 additional practice exams and answer keys

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## **Table of Contents Final Chemistry Exam Review Sheet With Answers**

1. Understanding the eBook Final Chemistry Exam Review Sheet With Answers
  - The Rise of Digital Reading Final Chemistry Exam Review Sheet With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Final Chemistry Exam Review Sheet With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Final Chemistry Exam Review Sheet With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Final Chemistry Exam Review Sheet With Answers



- Personalized Recommendations
  - Final Chemistry Exam Review Sheet With Answers User Reviews and Ratings
  - Final Chemistry Exam Review Sheet With Answers and Bestseller Lists
5. Accessing Final Chemistry Exam Review Sheet With Answers Free and Paid eBooks
- Final Chemistry Exam Review Sheet With Answers Public Domain eBooks
  - Final Chemistry Exam Review Sheet With Answers eBook Subscription Services
  - Final Chemistry Exam Review Sheet With Answers Budget-Friendly Options
6. Navigating Final Chemistry Exam Review Sheet With Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Final Chemistry Exam Review Sheet With Answers Compatibility with Devices
  - Final Chemistry Exam Review Sheet With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Final Chemistry Exam Review Sheet With Answers
  - Highlighting and Note-Taking Final Chemistry Exam Review Sheet With Answers
  - Interactive Elements Final Chemistry Exam Review Sheet With Answers
8. Staying Engaged with Final Chemistry Exam Review Sheet With Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Final Chemistry Exam Review Sheet With Answers
9. Balancing eBooks and Physical Books Final Chemistry Exam Review Sheet With Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Final Chemistry Exam Review Sheet With Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Final Chemistry Exam Review Sheet With Answers
- Setting Reading Goals Final Chemistry Exam Review Sheet With Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Final Chemistry Exam Review Sheet With Answers

- Fact-Checking eBook Content of Final Chemistry Exam Review Sheet With Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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