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final temperature is exactly halfway between  $t_1$  and  $t_2$  b the final temperature is exactly halfway between  $t_1$  and  $t_2$  the  
final temperature is closer to  $t_2$  than to  $t_1$  c

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water calculate the molar heat of combustion of hexane if the water temperature rises 55 4 C the specific heat capacity of  
water is 4 184 J/g C  $h_{\text{ms}}$   $t_{\text{h}}$  5 650 grams  $\text{H}_2\text{O}$  4 184 J/g C 55 4 C  $h$  1310 kJ

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ice at 34°C to steam at 138°C a sample of water absorbs  $1.41 \times 10^5 \text{ J}$  find the mass of the sample answers 1  $2.68 \times 10^4 \text{ J}$  2  $6.23 \times 10^4 \text{ J}$  3  $1.71 \times 10^4 \text{ J}$  4  $1.11 \times 10^5 \text{ J}$  5  $2.11 \times 10^5 \text{ J}$  6

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