

Class XI: Chemistry
Chapter 1: Some Basic Concepts of Chemistry

Top Formulae

Mass % of an element

1.
$$= \frac{\text{Mass of that element in the compound}}{\text{Molar mass of compound}} \times 100$$

2.
$$\text{Mass per cent} = \frac{\text{Mass of solute}}{\text{Mass of solution}} \times 100$$

3.
$$\text{Mole Fraction} = \frac{\text{No. of mole of a particular component}}{\text{Total No. of moles of solution}}$$

4.
$$\text{Molarity} = \frac{\text{No. of moles of solute}}{\text{Volume of solution in litres}}$$

5.
$$\text{Molality} = \frac{\text{No. of moles of solute}}{\text{Mass of solvent in kg}}$$

6.
$$\text{Moles of an element} = \frac{\text{Mass of element}}{\text{Atomic mass}}$$

7.
$$\text{Mass of one atom} = \frac{\text{Atomic mass}}{6.022 \times 10^{23}}$$

8.
$$\text{Moles of a compound} = \frac{\text{Mass of compound}}{\text{Molecular mass}}$$

9.
$$\text{Mass of one molecule} = \frac{\text{Molecular mass}}{6.022 \times 10^{23}}$$

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Rex M. Heyworth



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