

ORGANIC CHEMISTRY

FUNDAMENTALS

Types of Organic Compounds

Organic chemistry is the study of natural and synthetic materials that have carbon atoms as the key chemical feature. There are more than one million known organic compounds.

HYDROCARBON

- ALKANE**: $\text{C}_n\text{H}_{2n+2}$
- Ethane: C_2H_6
- Methyl Ethane: CH_3CH_3
- Ethyl Ethane: $\text{C}_2\text{H}_5\text{CH}_3$



ALKENE

- Ethene: C_2H_4
- Diene: Two $\text{C}=\text{C}$
- Triene: Three $\text{C}=\text{C}$



ALKYNE

- Ethyne: C_2H_2



AROMATIC

- Benzene: C_6H_6
- Aniline: $\text{C}_6\text{H}_5\text{NH}_2$



O- ADDED

ALCOHOL

- Methanol: $\text{Me}-\text{OH}$ (methyl alcohol)
- Phenol: $\text{Ar}-\text{OH}$
- Dialkylglycols: C_2-OHR
- Glycols: $\text{O}-\text{OHR}$



ETHER

- Ether: $\text{R}-\ddot{\text{O}}-\text{R}'$
- Diethyl Ether: Et_2O (diethyl ether)



EPoxy

- Epoxy: $\text{C}_2\text{H}_4\text{O}$ (ethylene oxide)



PEROXIDE

- Peroxide: $\text{R}-\ddot{\text{O}}-\text{O}-\text{R}'$
- Acetone Peroxide: $\text{C}_3\text{H}_5\text{O}_2$



BENZYL PERSCARATE: $\text{C}_6\text{H}_5\text{CH}_2\text{CO}_2$

>C=O ADDED

ALDEHYTE

- Methanal: H_2CO (formaldehyde)
- Benzaldehyde: $\text{Ar}-\text{CHO}$



KETONE

- 2-propanone: $\text{Me}_2\text{C}(=\text{O})\text{Me}$ (dimethyl ketone, acetone)
- Benzophenone: $\text{C}_6\text{H}_5\text{CO}-\text{C}_6\text{H}_5$



COO ADDITION

CARBOXYLIC ACID

- Ethanoic Acid: $\text{Me}-\text{COOH}$ (acetic acid)
- Acetate Ions: $\text{Me}-\text{COO}^-$
- Benzene Acids: $\text{Ar}-\text{COOH}$



ESTER

- Ethyl Acetate: $\text{Me}-\text{COO}-\text{CH}_2\text{CH}_3$



Other derivatives:

- Pronoyside: $\text{R}-\text{COO}-\text{OH}$
- Acid Anhydrides: $\text{R}-\text{COO}-\text{COR}'$

NITROGEN ADDED

AMINE

- Methyl Amine: $\text{H}_3\text{C}-\text{NH}_2$
- Phenylenamine: $\text{Ar}-\text{NH}_2$ (aniline)
- R-NH₂, $(\text{R})_2\text{NH}$ (R_3N)



NITRO

DIAZO

NITRILE

HALOGEN ADDED

HALOGEN

- Chloroform: $\text{Me}-\text{Cl}$ chloromethane
- Chlorobenzene: $\text{Ar}-\text{X}$
- Chloroethane: $\text{Ar}-\text{Cl}$
- Acyl Halides: $\text{R}-\text{CO}-\text{X}$
- Aryl Halides: $\text{Ar}-\text{X}$



Formulas & Isomers

- Molecular Formula:** Elemental symbols with subscripts denote the composition of a compound.
- Empirical Formula:** Subscripts denote the relative elemental composition.
- Chemical depiction:**
 - Dashed Formula:** Diagrams all atoms, show bonds as dashes.
 - Short-Hand Formula:** Hide H's, carbon atoms are depicted as lines, but other atoms are shown explicitly.



Newman projection formula



Dashed formula

- Newman Projections:** 2-d depiction.
- 3-dimensional: Wedges of sawhorse denote stereochemistry.

- Constitutional Isomers:** Different bonding connectivity (e.g., rings, bonds, branching, valence positions).
- Tautomer:** Easily interconverted structural isomers (e.g., keto-enol for ketone).
- Chirality:** Not identical with its mirror image.
- Antiplane:** Has a plane of symmetry (i.e., superimposable on its mirror image).
- Eplimers:** Pair of diastereomers which differ only in the configuration of one atom.
- More than 1 chiral center:**
 - n-chiral centers, $\leq 2^n$ stereoisomers.
 - Mixed 2 chiral centers, 4 isomers; 3 stereoisomers, 1 achiral (mirror-plane).

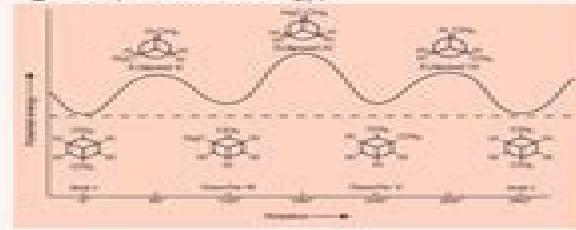


Arrows are clockwise

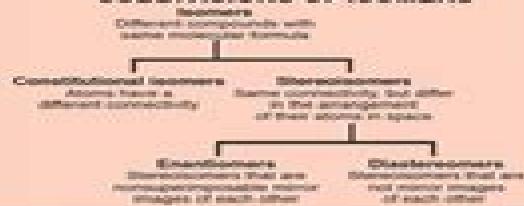
- Enantiomers:** 50/50 mixture of stereoisomers (no net optical activity).
- Stereoisomers:** Note R/S and (+/-) in the compound name (e.g., R (+) bromochloroethane).
- Fischer-Projections:** Diagram depicts chiral 3-D structures.
- Molecular Conformations:** Molecule exhibits structural variation due to free rotation about C-C single bond.



- Newman-Diagrams:** Depict rotation about a C-C bond: eclipsed (high energy), anti (low energy), and gauche (intermediate energy).



SUBDIVISIONS OF ISOMERS



that loses energy during the process of the reaction.

Heterocyclic: Non-carbon atom in the ring structure.

Hydrocarbon: Compound of H and C.

Ketone: Molecule made up of a monocyclic molecule bonded to a ketone.

that cannot be reduced by hydrolysis into another simple sugar.

Olefins: Alkenes.

Paraffinic: Alkanes.

Radioisotope: Isotope that is radioactive.

Saturated: Maximum # of H's (all C-C single bonds).

Tetrasaccharide: Monosaccharide with four carbon atoms.

Unsaturated: At least one C=C multiple bond.

Zymurgy: The study of the process of fermentation.

Common Terms

Aliphatic: Non-aromatic.

Aromatic: Benzene ring.

Conjugation: Sequence of alternating double (or triple) and single bonds.

Dipole-Dipole: Polar solvent stabilization.

Isomerization:

Exergonic Reaction: Chemical reaction

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