

Electroplating Lab

Data Table

Initial mass of copper electrode (g)	2.49 grams
Final mass of copper electrode (g)	2.36 grams
Initial mass of key (g)	7.48 grams
Final mass of key (g)	7.60 grams
Average current (A)	2.12
Time of current application (s)	100.8 seconds

Data Analysis (Part I)

1. Calculate the number of coulombs of charge that passed through the electrolytic cell. (Hint: make sure you are using the proper unit of time for your calculations)
2. Use your current vs. time data to calculate the theoretical number of moles of copper that should have plated out onto the brass key.
3. Use your data from measuring the mass of the electrode and the key to calculate the actual number of moles of copper that plated out.
4. Calculate the percent yield of copper.
5. Suggest the sources of error in your experiment.
6. Write the oxidation and reduction half-reactions for this process.

Post-Lab Questions

- 1) Turn in your data table. A blank has been provided

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Final mass of key (g)	7.60 grams
Average current (A) (calculated in question 4)	2.12
Time of current application (s)	100.8 seconds

Electroplating Lab Answers

Clifford Lane



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