

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (a) (a.i)-(a.ii) Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are consistent with experimental evidence. (b) (b.i)-(b.ii) Explain the relationship between atomic-scale phenomena and the physical theory describing the phenomena (atoms in terms of quantum mechanical calculations; readily accessible theory that can be used to explain all their mass, electrical charges, and locations within the atom. (c) (c.i)-(c.ii) Explain properties of atoms and their placement within periodic table in terms of configurations of their electrons, related to Pauli exclusion principle and arrangement of their four fundamental interactions (strong, electromagnetic, weak, gravity), such as their mass, charge, magnetic.

ANSWER KEY

Directions: Match the atomic theory selected on the correct card(s).

Model(s)	Model(s)	Atomic Theory's Basic Principles
(C)	Continuous (old MC)	A. Everything is made of atoms (that cannot be created, divided or destroyed)
(A)	Classical (old MC)	B. Matter (including atoms) consists mostly of positively charged nuclei, surrounded with negative distributed throughout the positive material
(B)	Quantum (new MC)	C. "matter is indivisible (unchangeable)"
(F)	Quantum (new MC)	D. Group Numbers (valence electrons)
(D)	Model (new MC)	E. Electron Cloud
(E)	Electrons (new MC)	F. Atoms are mostly empty space (the negatively charged electrons move)

Directions: Use the periodic table to answer the following questions:

- (C) = F, A, B** - The number of protons equals the number of electrons.
- atomic mass minus the atomic number will equal the number of neutrons.**
- (B) = (C)** - **(C) = (B)** - protons are positive (positively charged) and electrons are negative (negatively charged).

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JS Bruner



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