



Frame 9e Gas Turbine Manual

Amad Narto,Eha Vembriarta



Frame 9e Gas Turbine Manual:

Practicum Module Main Engine (Steam & Gas Turbine) Amad Narto, Eha Vembriarta, Steam Turbine and Gas Turbine are machineries used to drive a gear reduction which is connected to a generator and propeller shaft. The steam produced by the boiler steam boiler in a closed vessel is either a water pipe or a fire pipe. Steam pressurized is high enough flows through a nozzle which is directed to the turbine blades so that it rotates into another form of energy to produce powers. Boilers are designed to transfer heat from a combustion source which is usually fuel combustion. A boiler is a connected vessel in which water is filled and heated so that water boils and evaporates continuously into steam. Steam generated from the boiler is used to rotate the steam turbine blades and rotate the shaft axle which is connected to the electric generator. The propeller shaft rotates compressor driver cargo pump on tankers, Mooring Winch and Winch lass movers and others. As for the gas turbine, the combustion medium in the combustion chamber is gas. With a very high pressure it will rotate a low pressure turbine in which its rotation can be adjusted with a reduction gear and rotational speed regulator control which is often used for propulsion in aircraft and high speed ships. The operation of the Steam Turbine and Gas Turbine must comply with the operating standards that have been determined by regulations or safety laws. Therefore the user of the steam turbine and gas turbine will have to meet the standard of the manufacturer who already has a production permit. The standards made will ensure safety in operation thereby increasing the efficiency of the steam turbine while reducing the operating costs.

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the h^+ ions from the acid and the oh^- ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly
 exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va
 and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room
 temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable
 measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kj less
 exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
 sodium hydroxide solution with hcl as 57.9 kj mol⁻¹

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 produce salts neutralisation is the reaction between an acid and a base

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic
 eg 1 hcl aq + naoh aq → nacl aq + h₂o l for which Δh = 57.1 kJ mol⁻¹ eg 2 h₂so₄ aq + 2koh aq → k₂so₄ aq + 2h₂o l for which Δh = 57.1 kJ mol⁻¹

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$ you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts