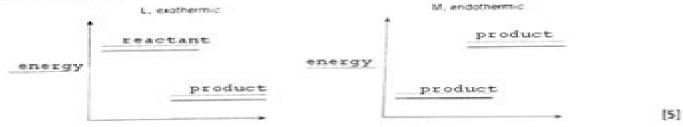
9.1 Energy changes in reactions

- 1. Describe these changes as either exothermic or endothermic.
 - a. The decomposition of copper(II) carbonate by heating Endothermic [1]
 - b. Burning paraffin Exothermic [1
 - c. Your tongue gets cold when you put sherbet on it. Exothermic [1]
 - d. Respiration in the cells of the body provides energy. Exothermic [1]
 - e. The temperature of the solution rises when concentrated hydrochloric acid is diluted.

endothermic	11	j

a. Complete these energy level diagrams for an exothermic and an endothermic reaction. Include an arrow in each diagram in the correct position.



Explain why energy level diagram L represents an exothermic reaction.

The reason is that in this reaction the energy or temperature of reactants

is higher than the product as during exothermic reaction is released [2]

Calcium carbonate can be converted to calcium oxide and carbon dioxide.

$$CaCO_{i}(s)$$
 — $CaO(s) + CO_{j}(g)$ Energy change, $\Delta H = +178$ kJ/mol

What type of energy change is taking place during this reaction and how does the information above show this?

Type of energy change endothermic [1]

How you know from the informatio Energy is positive in enthalpy change, which means the energy was taken in by the reactants.

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