

1. Describe these changes as either exothermic or endothermic.

- a. The decomposition of copper(II) carbonate by heating. Endothermic [1]
- b. Burning paraffin. Exothermic [1]
- c. Your tongue gets cold when you put sherbet on it. Exothermic [1]
- d. Respiration in the cells of the body provides energy. Exothermic [1]
- e. The temperature of the solution rises when concentrated hydrochloric acid is diluted.
..... endothermic [1]

2. a. Complete these energy level diagrams for an exothermic and an endothermic reaction. Include an arrow in each diagram in the correct position.



[5]

b. Explain why energy level diagram L represents an exothermic reaction.

The reason is that in this reaction the energy or temperature of reactants
.....
is higher than the product as during exothermic reaction is released. [2]

3. Calcium carbonate can be converted to calcium oxide and carbon dioxide.



What type of energy change is taking place during this reaction and how does the information above show this?

Type of energy change. endothermic [1]

How you know from the information: Energy is positive in enthalpy change, which means the energy was taken in by the reactants. [1]

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