



QUESTIONS

14. Answer these questions.

- 14.1 Which experiment would provide the most accurate measurement of the pH of a solution of ammonium nitrate?

- 14.1.1 Indirect titration
- 14.1.2 pH is [H⁺] = $\frac{1}{1000}$ M = 0.001 M
- 14.1.3 Universal indicator
- 14.1.4 pH is [H⁺] = 0.0001 M
- 14.1.5 pH measurement by titration

- 14.2 The concentration of a solution has been found to be 0.001 M. How many moles of solute are present in 100 mL of this solution? Please show any calculations clearly. Round your answer to the nearest gram. (3 marks) (4 marks)

1	2	3	4	5	6
7	8	9	10	11	12
13	14	15	16	17	18
19	20	21	22	23	24
25	26	27	28	29	30
31	32	33	34	35	36
37	38	39	40	41	42
43	44	45	46	47	48
49	50	51	52	53	54
55	56	57	58	59	60

- 14.3 Which of the following is the correct formula for the compound formed by the ions:

- 14.3.1 Ca^{2+} and Cl^{-}
- 14.3.2 CaCl
- 14.3.3 Ca_2Cl_2
- 14.3.4 Ca_2Cl

- 14.3.5 CaCl_2
- 14.3.6 Ca_2Cl_2

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