

- **Electron configuration:** Atomic orbital occupancy for an atom.
- **Pauli exclusion principle:** Each electron has a unique set of quantum numbers, an orbital may hold up to two electrons, one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$.
- **Hund's rule:** Electrons fill s , p , and d subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals), which is reinforced.
- **Ionization potential (IP):** Energy required to remove electron from an atom or ion.
 - **First IP:** Removal of first valence electron
- **Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; that scale, the Pauli scale, is:

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