

Stuff I Should Know (Page 2)

Complex Ions & Common Ligands

Ligands	polar molecules & anions	NH_3 , H_2O , OH^- , CN^- , Cl^-	Odd example: $\text{Fe}^{3+} + \text{SCN}^- \rightleftharpoons \text{FeSCN}^{2+}$
Central ions	transition metals and Al^{3+}	Ag^+ , Cu^{2+} , Ni^{2+} , Zn^{2+} , etc. & Al^{3+}	
Examples	usually twice the number of ligands as the charge on the central ion. Key Words: "excess, concentrated"	$\text{Ag}(\text{CN})_2^-$, $\text{Cu}(\text{NH}_3)_4^{2+}$, $\text{Ni}(\text{OH})_4^{2-}$, $\text{Zn}(\text{NH}_3)_4^{2+}$, $\text{Al}(\text{OH})_4^-$	Reaction with Acid: $\text{Cu}(\text{NH}_3)_4^{2+} + \text{H}^+ \rightarrow \text{Cu}^{2+} + \text{NH}_4^+$

Organic Chemistry & Functional Groups

alkanes $\text{C}_n\text{H}_{2n+2}$	alkenes C_nH_{2n}	alkynes $\text{C}_n\text{H}_{2n-2}$	aromatics (benzene) C_6H_6
alcohol $\text{R}-\text{OH}$	aldehyde $\text{R}-\overset{\text{O}}{\underset{\text{ }}{\text{C}}}-\text{H}$	ketone $\text{R}-\overset{\text{O}}{\underset{\text{ }}{\text{C}}}-\text{R}$	ether $\text{R}-\text{O}-\text{R}$
carboxylic acid $\text{R}-\overset{\text{O}}{\underset{\text{ }}{\text{C}}}-\text{OH}$	ester $\text{R}-\overset{\text{O}}{\underset{\text{ }}{\text{C}}}-\text{O}-\text{R}$	amine $\text{R}-\text{NH}_2$	amide $\text{R}-\overset{\text{O}}{\underset{\text{ }}{\text{C}}}-\text{NH}_2$
Substituted benzene:	ortho = 1,2	meta = 1,3	para = 1,4

nuclear chem

alpha ${}^4_2\text{He}$
betalelectron ${}^0_{-1}\text{e}$
neutron ${}^1_0\text{n}$
positron ${}^0_{+1}\text{e}$

ΔH ΔS Spont?

-	+	at all temps
+	+	high temps
-	-	low temps
+	-	no temps

Note: ΔS in J
 ΔG & ΔH in kJ

K_{sp} & Solubility, s

1:1	$K_{sp} = s^2$
1:2	$K_{sp} = 4s^3$
1:3	$K_{sp} = 27s^4$
2:3	$K_{sp} = 108s^5$

Lewis Acids & Bases

$\text{BF}_3 + \text{NH}_3 \rightarrow \text{BF}_3\text{NH}_3$
acid anhydrides (oxides of nonmetals, CO_2)
basic anhydrides (oxides of metals, MgO)
 $\text{MgO} + \text{CO}_2 \rightarrow \text{MgCO}_3$
decomposition reactions: $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
Strange Examples: $\text{P}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4$

Strange Ions: (nitride, N^{3-}) (hydride, H^-)



Flame Test Colors

Barium - green
Sodium - yellow
Copper - blue (w/ green)
Potassium - lavender
Strontium - red
Lithium - red
Calcium - orange

Quantum Numbers

n	1, 2, 3, ...
l	0 ... (n-1)
m	-l ... +l
m_l	+l, ..., -l
l	0 = s, 1 = p, 2 = d, 3 = f

Writing Lewis Structures

hint: use one valence electron to connect F's or Cl's then determine lone pairs (Ex: XeF_4)

Product-Favored (Spontaneous) Reactions

$\Delta G < 0$ $E^\circ > 0$ $K_{eq} > 1$

Properties Indicate Strength of Intermolecular Forces (IMF's)

IMF	BP	FP	H_{vap}	H_{fus}	VP
IMF	BP	FP	H_{vap}	H_{fus}	VP

Orders of Reactions & Graphs That Give Straight Lines

0 th Order	1 st Order	2 nd Order
$[\text{R}]$ vs. Time	$\ln[\text{R}]$ vs. Time	$1/[\text{R}]$ vs. Time
slope = -k	slope = -k	slope = -k

Electrochemical Cells

anode	cathode
oxidation	reduction
- side	+ side
lower E°	higher E°
e^- leave	e^- enter

Bond Orders

bond	B.O.	σ
single	1	σ
double	2	$\sigma + \pi$
triple	3	$\sigma + 2\pi$

SN & hybridization & shape

Steric Number	hybridization	basic shape
1	s	—
2	sp	linear
3	sp ²	Δ planar
4	sp ³	tetrahedral
5	sp ³ d	Δ bipyramidal
6	sp ³ d ²	octahedral

IMF's

London	nonpolar molecules, ex: CH_4 , He
dipole-dipole	polar molecules, ex: H_2S , SO_2
hydrogen bonding	H-F, H-O-, H-N-, NH_3 , H_2O amines and alcohols
metallic	metals, Ag, Pb
ionic	salts, NaCl , CaCO_3 (Note: "ates" contain covalent bonds)
covalent network	C (graphite), C (diamond), SiO_2 , WC, Si, SiC (Note: graphite = London, too)

Activity of Metals (Four Groups)

Metals	React with...
Groups I & II	H_2O ex: $\text{Li} + \text{H}_2\text{O} \rightarrow \text{Li}^+ + \text{OH}^- + \text{H}_2$
all others	Non-oxidizing Acid, ex: HCl $\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$
Cu, Ag, Hg	Oxidizing Acid, HNO_3 or H_2SO_4 (conc.) $\text{Cu} + \text{HNO}_3 \rightarrow \text{NO}_2 + \text{H}_2\text{O} + \text{Cu}^{2+}$
Au, Pt, Ir	Aqua Regia (HNO_3 + HCl)

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JS Bruner



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