

CHEM 121: Final Exam Study Guide

Chapter 1

- Know the Scientific method
- Know the definitions for hypothesis, scientific law, and scientific theory.

Chapter 2

- Length, mass, weight, volume
 - Know $1\text{ cm}^3=1\text{ mL}$ and $1\text{ dm}^3=1\text{ L}$
- Significant figures and Scientific notation
 - Rounding, in addition or subtraction, in multiplication or division
 - Keep as many sig figs until final answer
- Scientific notation
- Metric system
 - Know the metric prefixes: kilo (k), deci (d), centi (c), milli (m), and micro (μ)
 - Be able to perform metric-metric conversions using these prefixes.
- Use the metric-English conversions provided ($1\text{ in.}=2.54\text{ cm}$; $1\text{ lb}=453.6\text{ g}$; $1\text{ qt}=946\text{ mL}$)
- Volume by calculation
 - $V_{\text{rectangular solid}} = \text{length} \times \text{width} \times \text{thickness}$
- Volume by displacement
- Density: $d = \frac{m}{V}$
 - Calculate density, mass, or volume
 - Identify what items sink or float given densities of liquids and solids.
- Temperature
 - Know the formulas for converting "F-to-"C or "C-to-"F and K-to-"C or "C-to-K
- Percentage: ratio of parts per 100 parts
 - Given amount of part and whole, calculate %
 - Use a given % to solve for part or whole

Chapter 3

- Know that matter is studied at the macroscopic, microscopic, particulate (molecular) levels
- Physical states of matter
 - Determine physical state of substances (solids, liquids, gases) given descriptions of volume, shape, particles moving, etc.
- Classify properties and changes as physical or chemical
 - Know terms for changes of state:
 - Melting, freezing, vaporizing, condensation, sublimation, deposition
- Classification of matter
 - Given examples, determine which are elements, compounds, or mixtures
 - Given molecular-level images, determine which are elements, compounds, or mixtures and solids, liquids, or gases
 - Distinguish between homogeneous and heterogeneous mixtures
- Chemical reaction:
 - reactants: starting materials
 - products: substances produced in reaction

College Chemistry Final Exam Study Guide

Don E. Webster



College Chemistry Final Exam Study Guide:

College Chemistry Final Exam Study Guide Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**College Chemistry Final Exam Study Guide**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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