

Fifty Frequently Forgotten Facts of Chemistry

1. ATOMIC STRUCTURE & NUCLEAR CHEMISTRY

1) Protons are +1 each with a mass of 1 amu each, the number of protons = atomic number, nuclear charge = + (# protons). [Periodic Table]

- a) How many protons are there in a nucleus of Kr-85? _____
b) What is the nuclear charge of an atom of Br? _____
c) What is the mass of the protons in a nucleus of O-15? _____

2) Neutrons are neutral with a mass of 1 amu each, # neutrons = mass # - atomic number. Isotopes = atoms of the same element (same atomic #) but different # of neutrons (mass #). [Periodic Table]

- a) How many neutrons are there in the nucleus of $^{56}_{26}\text{Fe}$? _____
b) Circle the two nuclei that are isotopes of each other: $^{16}_8\text{O}$ $^{15}_7\text{N}$ $^{16}_8\text{O}$ $^{18}_9\text{F}$

3) Electrons are each -1 with a mass that is VERY, VERY tiny compared to the mass of a proton or neutron.

a) Which particle has a mass that is $1/1836^{th}$ the mass of a proton?

- 1) ^4_2He 2) ^1_1H 3) $^{0}_{-1}\text{e}$ 4) ^1_0n

4) Natural Decay: Parent Nuclide \rightarrow Decay particle + daughter nuclide [Tables N and O]

- a) Write the decay for U-238: _____
b) Write the decay for K-37: _____
c) Write the decay for P-32: _____

5) Artificial Transmutation is when a relatively stable nucleus is impacted by a particle bullet at high speeds and becomes an unstable nucleus of a different element. **Nuclear fission** occurs when nuclei of U-235 or Pu-239 are impacted by a neutron and split into two smaller nuclei and more neutrons. **Nuclear fusion** occurs when two small nuclei of hydrogen combine at high temperatures and pressures to form larger nuclei of helium. Both fission and fusion convert mass into a huge amount of energy.

Given the nuclear reactions:

- 1) $^{235}_{92}\text{U} + ^1_0\text{n} \rightarrow ^{92}_{36}\text{Kr} + ^{141}_{56}\text{Ba} + 3 ^1_0\text{n}$ 2) $^{239}_{94}\text{Pu} + ^4_2\text{He} \rightarrow ^{242}_{96}\text{Cm} + ^1_0\text{n}$
3) $_{81}^{208}\text{Po} \rightarrow _{80}^{208}\text{Bi} + ^1_0\text{e}$ 4) $^2_1\text{H} + ^2_1\text{H} \rightarrow ^4_2\text{He}$

- a) Which reaction represents natural decay? _____
b) Which reaction represents artificial transmutation? _____
c) Which reaction represents nuclear fission? _____
d) Which reaction represents nuclear fusion? _____

Fifty Forgotten Fun Facts Chemistry

Yan Bai



Fifty Forgotten Fun Facts Chemistry:

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