

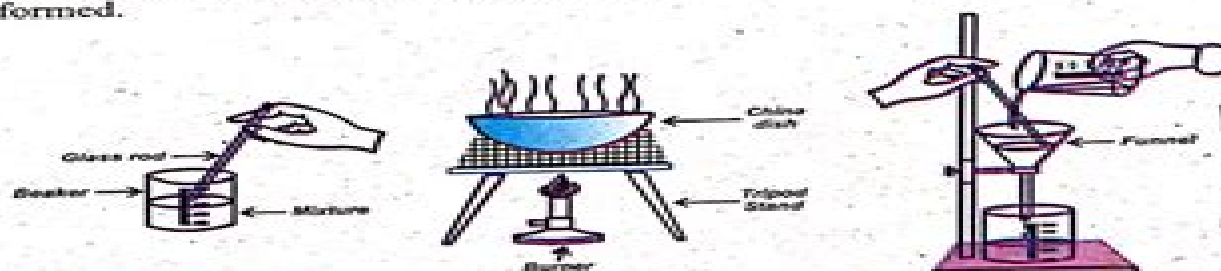
## CRYSTALLIZATION

# 1 EXPERIMENT

To prepare the crystals of Benzoic Acid from water.

### Principle:

Benzoic Acid is more soluble in hot water and very less (about  $.5\text{g/dm}^3$ ) soluble in cold water. Thus when hot saturated solution is cooled at room temperature solid crystals of benzoic acid are formed.



**Remember**  
Only 0.5g Benzoic acid is dissolved in  $1\text{dm}^3$   $\text{H}_2\text{O}$ .

### Apparatus:

Beaker, Glass rod, China dish, Stirrer, Filter paper, Tripod stand, Wire gauze

### Procedure:

Take about  $60\text{ cm}^3$  of distilled water in a  $250\text{ cm}^3$  beaker. Add some benzoic acid in it at room temperature. When no more acid dissolves then heat the solution on a tripod stand. Add further amount of acid in small fractions till a good quantity has been dissolved. Now filter the hot solution in china dish and cool it at room temperature. On cooling pure white lustrous flakes of crystals of benzoic acid are obtained. Decant the mother liquor. Dry the crystals between the folds of filter paper or by placing them in an oven or desiccator and show to the examiner.

### Precautions:

1. Don not prepare super saturated solution.
2. Always filter the hot solution
3. Do not give rapid cooling by placing it in ice bath.

### Result:

The pure crystals of benzoic acid are prepared and shown to the examiner.

**Remember**  
The melting point of pure Benzoic acid is  $121^\circ\text{C}$ .

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**American Council on Education**



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